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How To Enhance the Efficiency of Breslow Intermediates for SET Catalysis

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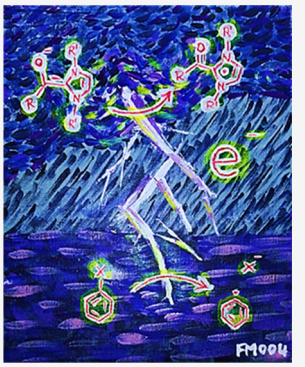
ABSTRACT: Oxidative carbene organocatalysis, which proceeds via single electron transfer (SET) pathways, has been limited by the moderately reducing properties of deprotonated Breslow intermediates **BI**$^-$s derived from thiazol-2-ylidene **1** and 1,2,4-triazolylidene **2**. Using computational methods, we assess the redox potentials of **BI**$^-$s based on ten different types of known stable carbenes and report our findings concerning the key parameters influencing the steps of the catalytic cycle. From the calculated values of the first oxidation potential of **BI**$^-$s derived from carbenes **1** to **10**, it appears that, apart from the diamidocarbene **7**, all the others are more reducing than thiazol-2-ylidene **1** and the 1,2,4-triazolylidene **2**. We observed that while the reducing power of **BI**$^-$s significantly decreases with increasing solvent polarity, the redox potential of the oxidant can increase at a greater rate, thus facilitating the reaction. The cation, associated with the base, also plays an important role when a nonpolar solvent is used; large and weakly coordinating cations such as Cs^+ are beneficial. The radical−radical coupling step is probably the most challenging step due to both electronic and steric constraints. Based on our results, we predict that mesoionic carbene **3** and abnormal NHC **4** are the most promising candidates for oxidative carbene organocatalysis.



INTRODUCTION

While stable singlet carbenes have found numerous applications when associated with metals,[1,2] they also display a rich chemical reactivity on their own merits. Thiazol-2-ylidenes **1**[3] and 1,2,4-triazolylidenes **2**[4] have been long known to induce umpolung reactivity of carbonyl compounds, giving rise to the formation of nucleophilic Breslow intermediates (**BI**s).[5,6] The latter allow for a variety of chemical transformations, which proceed in a well understood ionic mode, via electron-pair-transfer mechanisms. Several decades ago, it was shown that **BI**s are also involved in single-electron transfer (SET)-based catalysis during the oxidative decarboxylation of pyruvate to form acetyl-CoA.[7] However, it is only in 2008 that Studer and co-workers[8] developed a TEMPO-mediated biomimetic oxidation of aldehydes to TEMPO-esters. Since that time, other oxidants, such as nitroarenes, nitroalkenes, CX_4, C_2Cl_6, sulfonic carbamate, redox-active esters, the Togni's reagent, polyfluoroalkyl halides, Katritzky pyridinium salts, oxime ester, and recently aryl iodides (vide infra), were also employed to achieve oxidative reactions with aldehydes.[9]

It was initially postulated that the catalytic cycle proceeded through a SET from the **BI**s leading to the radical cations **BI**$^{\bullet+}$ which was regarded as the key intermediates. However, mechanistic and electrochemical investigations of standard enols,[10] as well as recent studies, showed that the active paramagnetic species is rather the corresponding neutral radical **BI**$^\bullet$.[11] Moreover, we showed that the SET does not occur from the **BI**, but from their deprotonated form, **BI**$^-$.[11] Consequently, the catalytic cycle of the oxidative reactions of aldehydes involves six steps, which are described in Figure 1.

So far, with the exception of our work[12] employing mesoionic carbene (MICs) **3**,[13] thiazol-2-ylidenes **1**[3] and 1,2,4-triazolylidenes **2**[4] have been the only carbenes efficiently used in the catalytic oxidative reactions of aldehydes. Due to the moderately reducing properties of **BI**$^-$s derived from **1** and **2**, the reported SET catalyzed reactions require a relatively strong oxidant (*vide supra*), which dramatically limits their synthetic applications. Motivated by the availability of a library of stable singlet carbenes with varying electronic and steric properties, we wondered if we could predict the best candidates to promote these reactions. Obviously, one of the key factors is the reducing power of the **BI**$^-$. Since electrochemical studies are difficult to implement for highly reactive intermediates such as Breslow enolates, we used computational methods to assess the redox potentials of **BI**$^-$s based on a variety of known stable carbenes. We also report our findings concerning each step of the catalytic cycle, including the possible pitfalls.

DFT calculations[14] (in Jaguar 9.1)[15] at the PBE/6-31G**//cc-pVTZ[16] level of theory with the Poisson−Boltzmann solvation model were used for carbenes **1**−**10** (Figure 2) and their related catalytic intermediates. All molecules and ions

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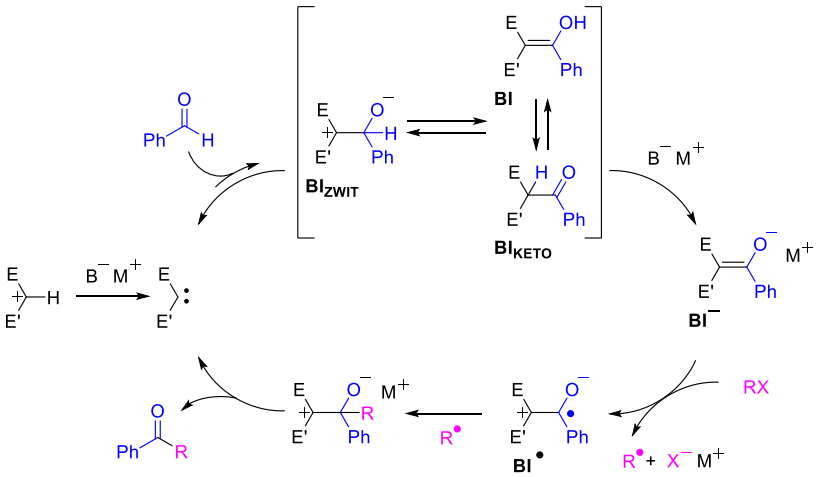

Figure 1. Catalytic cycle for carbene oxidative organocatalytic reactions.

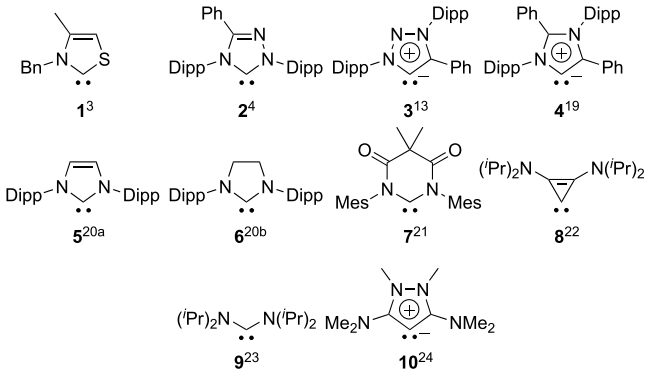

Figure 2. Carbenes considered in this manuscript.

were considered as isolated structures. All redox potentials were computed from solution free energies and referenced to the computed ferrocene/ferrocenium couple.[17] Values are given in acetonitrile unless noted otherwise. To validate our choice of computational methodology, an array of basis set/functionals were evaluated (see Supporting Information (SI)) and the results were compared to known experimental values. Comparison to ten literature-known redox potentials showed an average error of 190 mV with five examples showing an error of 100 mV or less which is in the range of experimental error. Larger errors may indicate the significance of aggregates (contact/solvent-bridged/solvent-separated ion pairs)[18] in some cases under the respective experimental conditions.

RESULTS AND DISCUSSION

First and Second Oxidation of BI$^-$s. We found that the first (E^1) and second (E^2) oxidation potentials of **BI$^-$**s based on carbenes **1**–**10** and benzaldehyde span over a large range from −1.50 (**2**) to −2.27 V (**5**), and −0.31 (**7**) to −2.43 V (**10**), respectively (Table 1 and Figure 3).

Because of the rather high computational expense of modeling redox potentials, we also looked into simpler predictive indicators connecting the carbenes' electronic properties and the redox potential of their **BI$^-$**s. The HOMO orbital energies of carbenes **1**–**10** (and additional examples detailed in the SI) give a rough indication of the redox potential ranking of **BI$^-$**s (Pearson-*r*: −0.666), but notable outliers exist. For example, the computed redox potentials for the **BI$^-$**s derived from MIC **3** (−2.22 V) and

Table 1. HOMO and LUMO Energies of Carbenes 1–10 and Redox Potentials of the Corresponding BI$^-$s and BIs

	HOMO[a]	LUMO[a]	E^{1b}	E^{2b}	$E^{red b}_{BI}$	$E^{ox a}_{BI}$
1	−5.05	−1.51	−1.58	−0.72	−1.57	−0.57
2	−5.22	−2.12	−1.50	−0.93	−1.65	−0.61
3	−4.65	−2.14	−2.22	−1.12	−1.94	−0.95
4	−4.29	−2.18	−2.16	−1.59	−2.00	−1.40
5	−4.94	−1.23	−2.27	−1.10	−2.10	−0.88
6	−4.78	−1.12	−1.94	−1.05	−1.97	−0.57
7	−4.99	−2.81	−1.05	−0.31	−1.65	+0.37
8	−4.20	−0.15	−2.26	−1.43	−2.47	−1.06
9	−3.63	+0.15	−1.89	−1.14	−2.16	−0.58
10	−3.59	−0.87	−2.25[c]	−2.43[c]	−2.34	−1.91

[a]In eV. [b]In V. [c]See SI for more details.

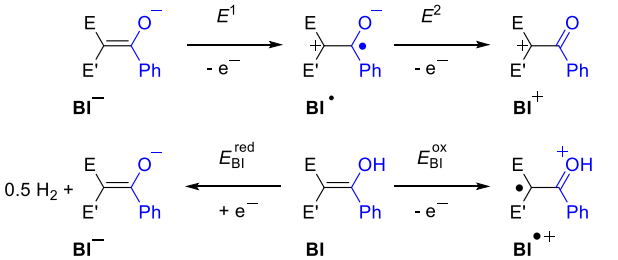

Figure 3. First and second oxidation of **BI$^-$** (top); reduction and oxidation of **BI** (bottom).

imidazol-2-ylidene **5** (−2.27 V) are predicted to be comparable, while their HOMO levels are significantly different (−4.65 and −4.94 eV, respectively). Just as for organometallic complexes, the π-accepting ability of carbenes plays a significant role. Thus, a better predictor can easily be obtained using $\frac{1}{2}(\varepsilon_3^{HOMO} + \varepsilon_3^{LUMO})$ which is closely related to Mulliken's electronegativity (Pearson-*r*: − 0.74, R^2: 0.55, Figure 4).[25] Indeed, **3-BI$^-$** exhibits a lower LUMO than **5-BI$^-$** (LUMO: − 2.14 and −1.23 eV, respectively) which readily explains that despite its higher HOMO, the reduction power of its **BI$^-$** is comparable to that of **5**.

From these data, it clearly appears that MICs **3**[13] and *a*NHCs **4**,[19] which lead to highly reducing **BI$^-$**s, are more promising candidates than thiazol-2-ylidenes **1**[3] and the 1,2,4-triazolylidenes[4] that have widely been used for SET catalysis.

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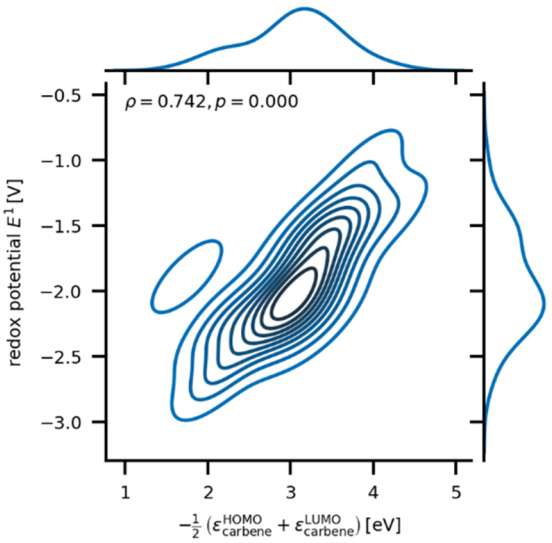

Figure 4. Distribution plot showing the correlation between E^1 and $\frac{1}{2}(\varepsilon_3^{HOMO} + \varepsilon_3^{LUMO})$ for all computed carbene/**BI**$^-$ pairs, see SI for a full list.

Their strongly negative redox potential indicates that they can react with weak oxidants.

Oxidation and Reduction of BIs. Because of the earlier assumption that **BI**s were the active reducing agent, we also calculated their oxidation and reduction potentials. We found that the oxidation of **BI**s also covers a large range [+0.37 (**7**) to −1.91 V (**10**)] and roughly follows the order observed for **BI**$^-$s. Importantly, they are on average 1.1 V less reducing than the corresponding **BI**$^-$s. The resulting **BI**$^{\bullet+}$s are strongly acidic as shown by the predicted large negative pK_as (see SI), which readily explain the involvement of the neutral radicals **BI**$^\bullet$ instead of the radical cations **BI**$^{\bullet+}$ in catalysis.[11]

A reduction process was also experimentally observed by cyclic voltammetry for a Breslow homoenol derived from **6**, in the form of an irreversible peak at −1.90 V. The outcome of this process was unclear.[11] Our calculations show that the reduction of the **BI** derived from **6**, computed at −1.97 V, is coupled with a fast or concerted hydrogen evolution resulting in the formation of the corresponding **BI**$^-$ (see SI). Interestingly, with some carbenes (**1**, **3**−**5**), the **BI**$^-$ can readily reduce the corresponding **BI** giving **BI**$^\bullet$, half an equivalent of H_2, and regenerating the **BI**$^-$. The overall process is the electrocatalytic reduction of **BI**s into **BI**$^\bullet$s and H_2. In these cases, this competing pathway could prevent efficient oxidative catalytic processes.

Influence of the Solvent and Cation on the Reducing Properties of BI$^-$s. After examining the role of the carbene, we turned our attention to the influence of the solvent and cation on the reducing properties of **BI**$^-$s. We modeled solvation with a self-consistent reaction field approach (SCRF) and calculated the solvation with 25 distinct dielectric constants in the range of 1 (vacuum) to 100 (Figure 5). Note that this model does not capture dispersion effects or explicit coordination which implies that apolar solvents may not be captured accurately. We observed that the reductant strength of the **BI**$^-$ significantly decreases with increasing polarity of the solvent, which is a stabilizing factor for charged molecules. For example, the redox potential of the couples **BI**$^-$/**BI**$^\bullet$ derived from **1** and **3** shifts from −2.42 and −2.92 V in *t*BuOMe (ε = 4.5) to −1.48 and −2.15 V in water (ε = 78.4). Therefore, at first glance, we could assume that oxidative catalysis involving **BI**$^-$s should be favored by nonpolar solvents, but this is not necessarily the case, as evidenced by recent work by Ohmiya and co-workers.[26] They reported the arylacylation of styrene in DMSO/H_2O, with phenyl iodide as an oxidant (E_{pa} = −2.69 V in CH_3CN), and thiazol-2-ylidene **1** as a catalyst, the **BI**$^-$ of which has a reported $E_{1/2}$ = −1.43 V in CH_3CN. To rationalize this thermodynamically unfavorable reduction, they cite the previous work by Saveant and co-workers[27] and wrote "the small reorganization energy of the enolate form of the Breslow intermediate and the fast mesolytic cleavage of the C(sp^2)−I bond makes the pathway kinetically feasible". Interestingly, our calculations show that although the redox potentials of **BI**$^-$s increase with solvent polarity, the redox potential of PhI (PhI → Ph$^\bullet$ + I-) increases at a higher rate. This is mainly due to the larger ion size of the **BI**$^-$ compared to I-; the latter being better stabilized by polar solvents. Consequently, the reaction becomes thermodynamically favorable, as shown by ΔG which decreases from −0.71 in *t*BuOMe to −3.81 kcal·mol^{-1} in water for **1** (Figure 5, left). In the case of carbene **3**, the ΔG changes from −12.63 in

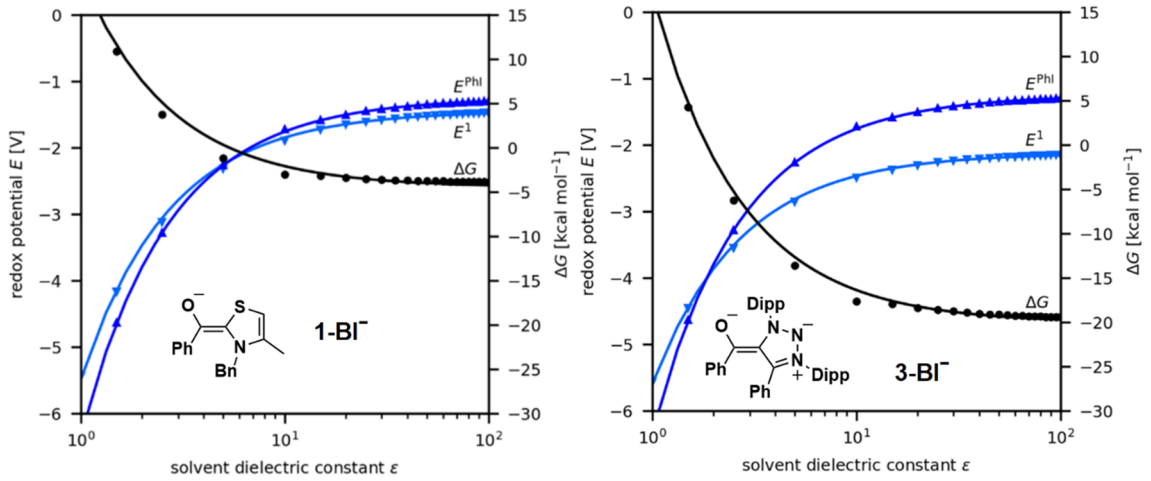

Figure 5. Thermodynamic feasibility of PhI reduction by **BI**$^-$ responds differently with different carbenes which is illustrated for **1** (left) and **3** (right) with the corresponding redox potentials and the redox reaction's ΔG.

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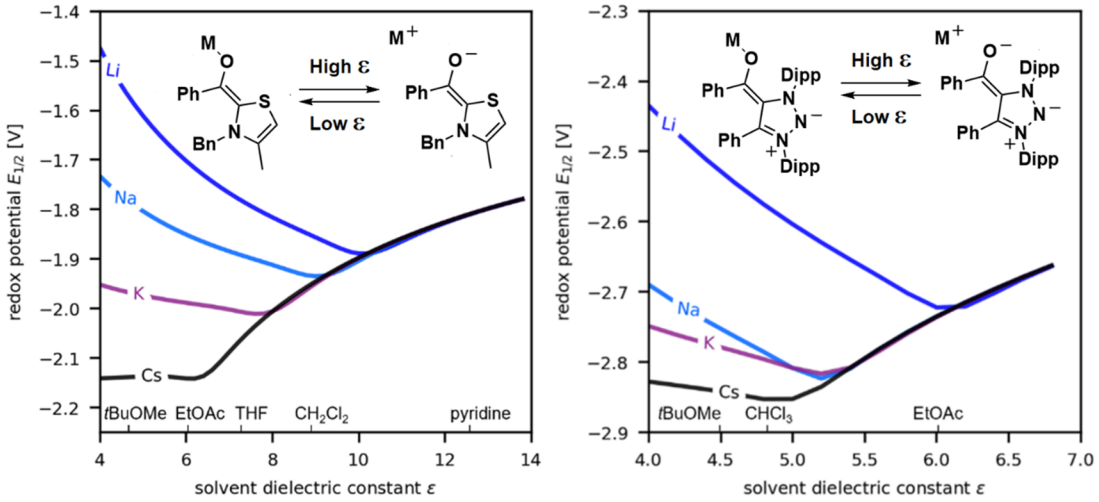

Figure 6. Influence of the cation (Li, N, K, Cs) on the first reduction potential of $\mathbf{BI}^-\mathbf{M}^+$ salts derived from carbenes **1** (left) and **3** (right) with the solvent polarity.

*t*BuOMe to −19.35 kcal·mol^{-1} in water (Figure 5, right). Due to the higher reducing power of MIC **3** compared to thiazol-2-ylidene **1**, it is not surprising that the former is able to promote the arylacylation of styrene in *t*BuOMe under milder conditions[12] than those reported for thiazol-2-ylidene **1** in water.[26]

Regarding the role of the cation, we limited our study to alkali metals (Li, Na, K, and Cs) because of their predominant use as a countercation of the base in oxidative organocatalysis. The salts were modeled, and a Boltzmann two-state distribution was used to predict the respective equilibrium position of the association/dissociation ($\mathbf{BI}^-\mathrm{M}^+ \rightleftarrows \mathbf{BI}^- + \mathrm{M}^+$). This model ignores distinct states mediating between isolated and coordinated/contact ion pairs. Particularly the transition area may behave differently due to the formation of solvent-separated or solvent-bridged ion pairs.[18] We used the $1/\varepsilon$ dependency of the employed solvation model to inter- and extrapolate from three explicitly computed solvation energies (in CH_3CN ε = 37.5, THF ε = 7.6, and *t*BuOMe ε = 4.5). Unsurprisingly, we found that, in polar solvents, the role of the cation is minimal due to the solvent-induced $\mathbf{BI}^-/\mathrm{M}^+$ separation. In contrast, in nonpolar solvents, the smaller cations significantly increase the redox potential of $\mathbf{BI}^-\mathrm{M}^+$ (Figure 6). Therefore, large and weakly coordinating cations are beneficial in a nonpolar medium.

Formation of the Deprotonated Breslow Intermediate ($\mathbf{BI}^-$). The initial step of the catalytic cycle in carbene oxidative organocatalysis is the deprotonation of the carbene-conjugate acid. Our results, combined with those found in the literature, suggest that even carbonates are basic enough for all carbene precursors, provided that an aldehyde is present to shift the equilibrium.[26]

The carbene must be nucleophilic enough to react with benzaldehyde in order to obtain the primary form zwitterionic adduct $\mathbf{BI}_{\mathrm{ZWIT}}$. We found this step to be endergonic with all carbenes **1**−**10** (ΔGs ranging from +3.7 to +18.2 kcal·mol^{-1}), but if we combine this process with the tautomerization into **BI** or $\mathbf{BI}_{\mathrm{KETO}}$, the free enthalpies range from −15.0 (**7**) to +11.6 kcal·mol^{-1} (**10**) (Table S9), suggesting that this chemical transformation should be achievable at room temperature with all carbenes **1**−**10**. Note that the $\mathbf{BI}_{\mathrm{KETO}}$ tautomer is rarely considered, although it has been observed experimentally. Berkessel and co-workers[28] have reported the rearrangement of **BI** to $\mathbf{BI}_{\mathrm{KETO}}$ with NHCs of types **5** and **6**, and our group has shown that the $\mathbf{BI}_{\mathrm{KETO}}$ tautomer was the thermodynamic product when cyclic (alkyl)(amino)carbenes[29] were reacted with benzaldehyde.[30]

We calculated the ΔG for the deprotonation of any tautomer of Breslow intermediates using ^{t}BuOK as a base and acetonitrile as the solvent. This process is exergonic for carbenes **1**−**10** (up to ΔG = −24.7 kcal·mol^{-1} for thiazol-2-ylidene **1**) (Table S6).

Coupling of $\mathbf{BI}^\bullet$ with Organic Radicals. It is well understood that radical coupling is thermodynamically favorable when both radicals have similar singly occupied molecular orbital (SOMO) energies. We found that the SOMO levels of $\mathbf{BI}^\bullet$s generated from carbenes **1**−**10** lie between −4.06 (**7**) and −2.65 eV (**8**), and therefore they are electron-rich radicals (Figure 7). The $\mathbf{BI}^\bullet$s derived from the most exploited thiazol-2-ylidene **1** and 1,2,4-triazolylidene **2** are among the least electron-rich of the series (−3.43 and −3.29 eV, respectively). Those derived from carbenes **3**−**5**, **8**, and **9** are the most electron-rich examples (>−2.96 eV). The SOMO energy level of $\mathbf{BI}^\bullet$s readily explains the type of coupling partners which have been successfully and unsuccessfully used experimentally thus far. Tertiary alkyl radicals (^{t}Bu: − 3.66 eV) work very well, but phenyl radicals are far too electron-poor (−5.53 eV) to directly couple with $\mathbf{BI}^\bullet$s. However, $Ph^\bullet$ can add to carbon−carbon double bonds, such as in styrene, generating a rich benzylic 1,2-diphenylethyl radical (−4.12 eV), which can couple with $\mathbf{BI}^\bullet$s. This is the concept of radical relay, which has been successfully used experimentally.[12,26,9d] Based on the calculated SOMO energy levels, further promising candidates are $i\mathrm{Pr}^\bullet$ (−3.99 eV), benzyl$^\bullet$ (−4.35 eV), and radicals in α-position of an heteroatom lone pair such as $\mathrm{EtOCH^\bullet CH_3}$ (−3.31 eV). Note also that it is easy to tune the redox potential of some of the $\mathbf{BI}^\bullet$s, especially those for which an electron-withdrawing group can be placed in the α-position of the carbene center. For example, in the MIC-derived $\mathbf{BI}^\bullet$s series, the replacement of the phenyl group on the carbon atom of **3** by COOMe results in the decrease of the SOMO energy level from −2.85 eV to −3.23 eV.

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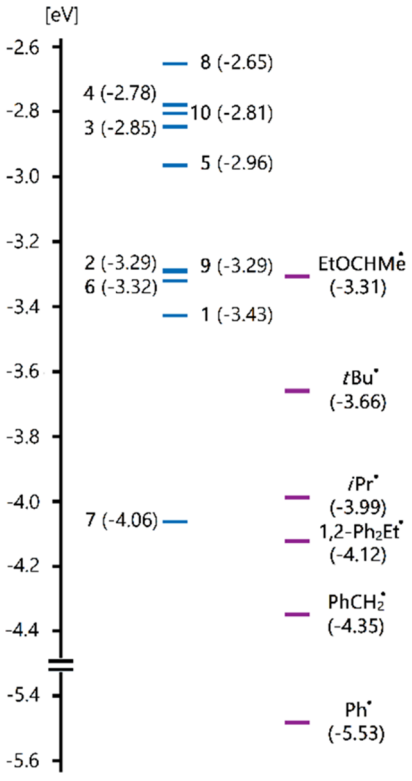

Figure 7. SOMO energy levels of **BI•** derived from carbenes **1**−**10**.

The radical−radical coupling step is also probably the most challenging step in carbene oxidative organocatalysis from a steric standpoint. For example, in the MIC-catalyzed arylacylation of styrenes, the replacement of an H by a phenyl group on the carbon next to the carbene center almost entirely suppressed the reaction.[12] Outside of their early discovery and affordable production, a major reason for the success of NHCs like **5** and **6** as ligands in transition metal catalysis is their large steric bulk which improves the reductive elimination step. In contrast, in carbene oxidative organocatalysis, steric bulk hampers the efficiency of the radical−radical coupling step, which can explain the superiority of **1** and **3**. To evaluate the steric demands associated with the **BI•**s, we computed their buried volumes (V_{bur}),[31] a technique usually applied to NHC transition metal complexes.[32] The V_{bur} was measured within a sphere of a 3.5 Å radius centered at the carbonyl carbon. Topographic steric maps show the large difference between **BI•**s derived from thiazol-2-ylidene **1** (**1-BI•**) and imidazolylidene **5** (**5-BI•**) in the amount of steric protection (Figure 8). Carbenes known to promote oxidative organocatalytic reactions lead to **BI•** with buried volumes below 70%. In the case of **3**, substituting the Dipp groups on N atoms with Ph groups and the Ph on the carbon atom with an H atom decreases the V_{bur} from 66% to 52%. In this context, *a*NHC **4**, which is also stable with a hydrogen on the alpha-carbon seems particularly promising, yielding a V_{bur} of 54% with Ph groups in all the other positions.

All carbenes **1**−**10** are nucleophilic enough to react with benzaldehyde to give the corresponding BIs. However, according to the HOMO energies, the typically used carbenes **1** and **2** are the least nucleophilic of the series, and thus the other carbenes could possibly allow for the use of less electrophilic partners than aryl aldehydes. The deprotonation of the BIs is exergonic in all cases, and thus is not a hurdle.

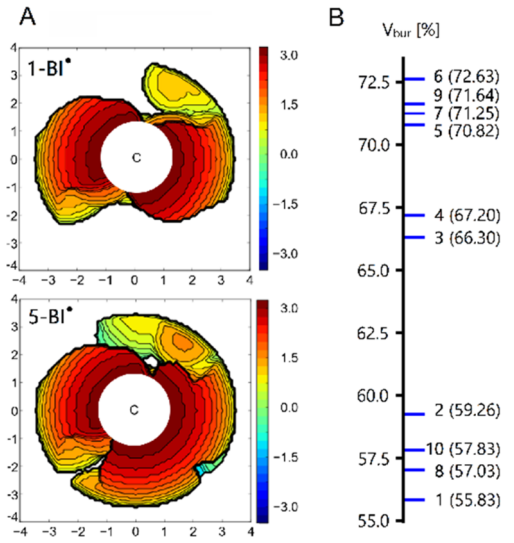

Figure 8. (A) Steric maps of **1-BI•** and **5-BI•**. (B) Buried volumes of **BI•** derived from carbenes **1**−**10**.

From the calculated values of the first oxidation potential of BI^- derived from carbenes **1**−**10**, it clearly appears that, apart from the diamidocarbene **7**, all the others are more reducing than thiazol-2-ylidene **1** and the 1,2,4-triazolylidene **2**. We observed that the reducing power of BI^- significantly decreases with increasing polarity of the solvent. At first glance, this could imply that oxidative catalysis should be favored by nonpolar solvents, but we found, with PhI as an example, that the redox potential of the oxidant can increase at a higher rate, and thus can facilitate the reaction. The cation, associated with the base used to deprotonate the conjugate acid of the carbenes, also plays an important role when a nonpolar solvent is used; large and weakly coordinating cations such as Cs+ are beneficial. The radical−radical coupling step is probably the most challenging step in carbene oxidative organocatalysis due to both electronic and steric constraints. According to the calculated SOMO energy level, all the BI•s derived from carbenes **1**−**10** are electron-rich, which readily explains the type of coupling partners which have been successfully and unsuccessfully used so far experimentally. Among radicals which have not yet been used, we found ^{i}Pr•, tolyl•, and radicals in the *α*-position of a heteroatom lone pair should work. Importantly, the range of the promising radical candidates could be expanded by tuning the redox potential of BI•s. This can be readily accomplished for carbenes in which an electron-withdrawing group can be placed in the *α*-position of the carbene center; MIC **3** and aNHC **4** are excellent candidates. Steric bulk can also hamper the efficiency of the radical−radical coupling step, which can explain the superiority of **1** and **3** over imidazol-2-ylidenes **5** and imidazolin-2-ylidenes **6**, which require bulky substituents on both nitrogen atoms for their stability. In this context, MIC **3** and aNHC **4**, which are stable with a hydrogen on the carbon *α* to the carbene center, appear particularly promising.

EXPERIMENTAL SECTION

Theoretical procedures and details are included in the Supporting Information. This material is available free of charge via the Internet at http://pubs.acs.org. The code for some of the data exploration and for making some of the figures can be found at: https://github.com/BaikgrpKAIST/FIGS-BIs_for_SET_catalysis.

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■ ASSOCIATED CONTENT

Data Availability Statement

The data underlying this study are available in the published article, in its Supporting Information, and openly available in GitHub at https://urldefense.com/v3/__https://github.com/BaikgrpKAIST/FIGS-BIs_for_SET_catalysis__;!!Mih3wA!C8zqV9pGFD-ZkMn97f8lkay3DB_iJcbkEpBCBAbrelAM8TgSuZePRuDJb0i743Jh9R2-puj_v-ZdDRMgawezxo9kzlOAGJc$.

Supporting Information

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Theoretical procedures and details (PDF)

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Author Contributions

Calculations were performed by F.F.M. All authors conceived and designed the project, analyzed the data and wrote the manuscript.

Notes

The authors declare no competing financial interest.

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